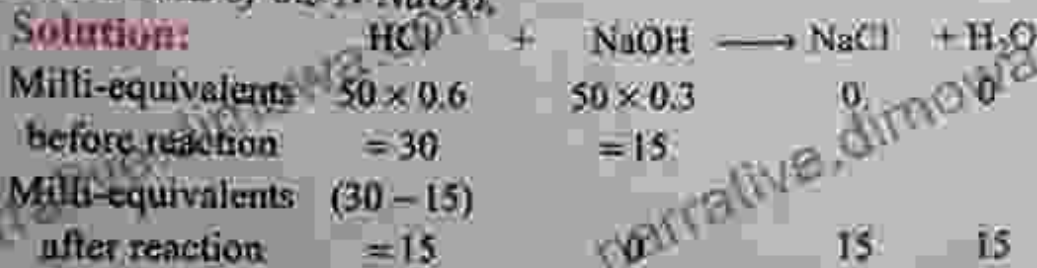


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Example 7. Calculate the $[Cl^-]$, $[Na^+]$, $[H^+]$, $[OH^-]$ and pH of the resulting solution obtained by mixing 50 mL of 0.6 N HCl and 50 mL of 0.3 N NaOH.



Total volume = 50 + 50 = 100 mL.

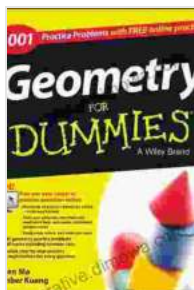
$$[Cl^-] = \frac{15 + 15}{100} = 0.3 M$$

$$[Na^+] = \frac{15}{100} = 0.15 M$$

$$[H^+] = \frac{15}{100} = 0.15 M$$

$$[OH^-] = \frac{10^{-14}}{0.15} = 6.6 \times 10^{-14}$$

$$pH = -\log [H^+] = -\log 0.15 = 0.8239$$



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★★★★☆ 4.5 out of 5

Language : English
 File size : 15526 KB
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 Screen Reader : Supported
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Covering a vast array of chemistry topics, "Chemistry 1001 Practice Problems For Dummies" provides a comprehensive review of:

- Atomic structure and bonding
- Chemical reactions and stoichiometry
- Gases, liquids, and solutions

li>Thermochemistry

- Equilibrium and kinetics
- Acids, bases, and salts
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- Organic chemistry

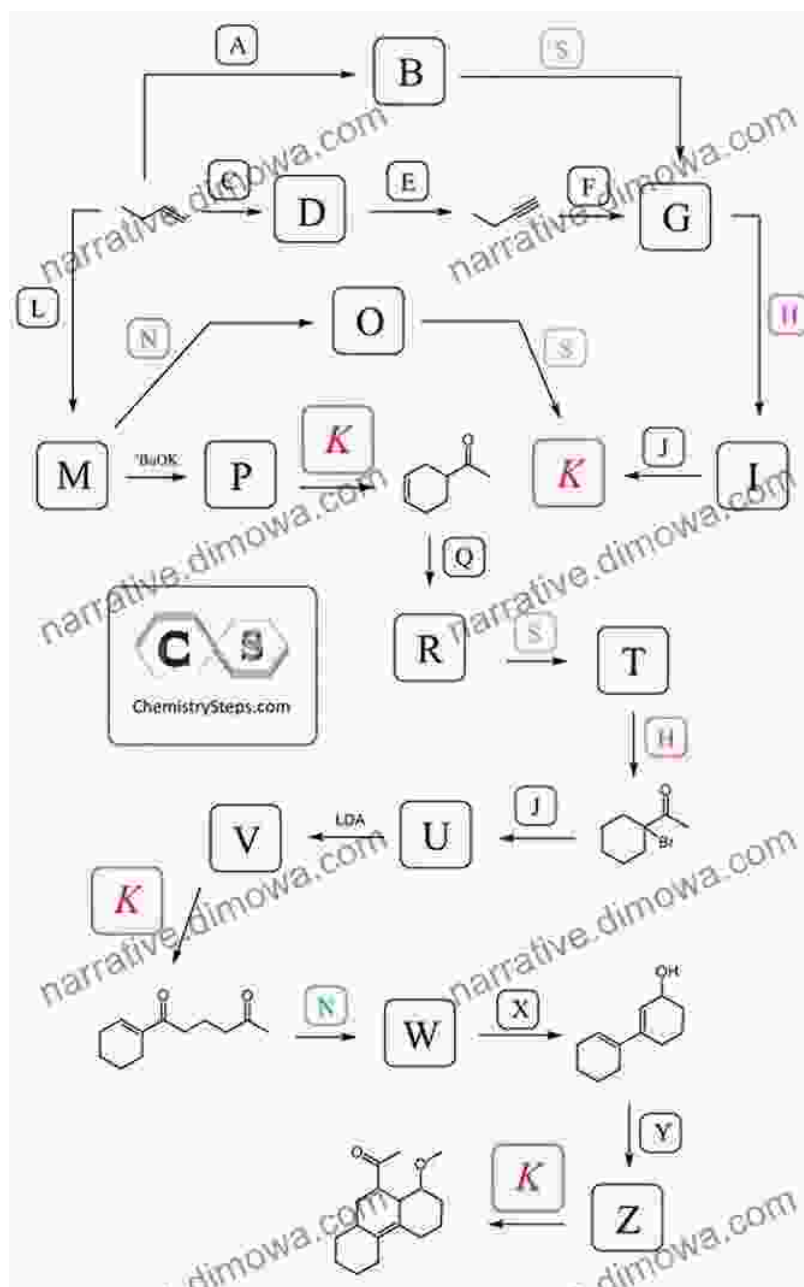
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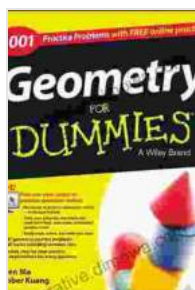
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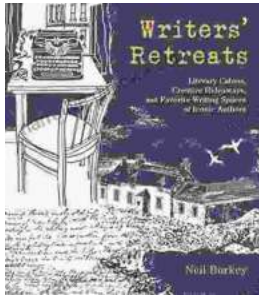


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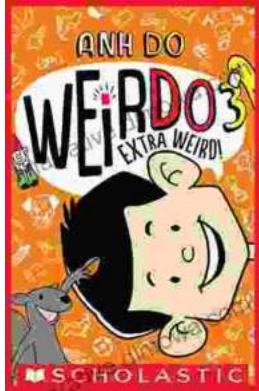
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